



$$\text{Molar ratio} = \frac{\left[\frac{I_{cpd}}{nH_{cpd}} \right]}{\left[\frac{I_{std}}{nH_{std}} \right]}$$

$$\text{wt}\% = \frac{mg_{std} \times MW_{cpd} \times \text{molar ratio} \times P_{std}}{mg_{cpd} \times MW_{std}} \times 100$$

Internal standard = 88 mg
 Sample = 142 mg
 Molar ratio = $[0.90/2]/[1.00/2] = 0.90$ wt% = 96%
 Molar ratio = $[2.74/6]/[1.00/2] = 0.91$ wt% = 97%
 %P of standard = 1
 MW of Product = 210.27
 MW of standard = 122.12
 wt% = 97%

